

Unit
3

Unit
4

Unit
5

STAND

Lesson 91 Nomenclature Type I

Text 4.1

Group 1, 2 metals are monatomic
positive charge is equal to group number

Groups 5A, 6A, 7A charge: $(\text{group} - 8) = \text{charge}$

Monatomic cations are identified by element's name

Examples for cations

Na sodium

Na⁺ sodium ion (or cation)

Ca calcium

Ca²⁺ calcium ion

Monatomic anions drop the ending of element name

Examples for anions

Cl chlorine

Cl⁻ chloride

S sulfur

S²⁻ sulfide

N nitrogen

N³⁻ nitride

↓
△ ide ending

Note 4A + 8A usually do not form ions.